

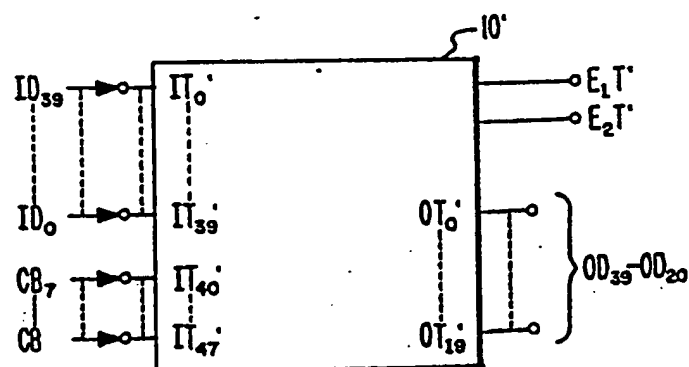
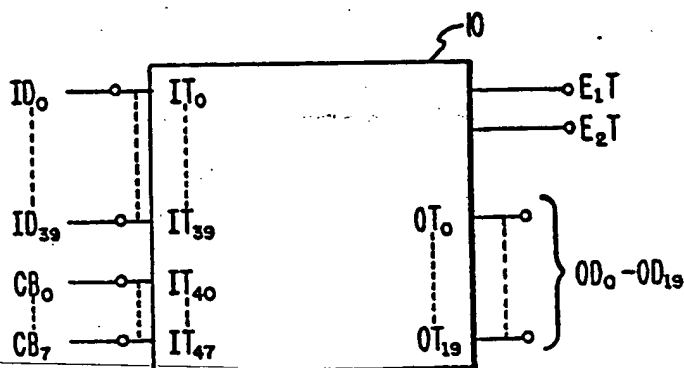


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(54) Title: LANGMUIR BLODGETT FILMS**(57) Abstract**

Novel Langmuir-Blodgett (LB) films of ABAB structure in which the compounds A and B are different and contain complementary groups which result in non-centrosymmetry of the layer plane. Three types of structure are described (i) in which A and B have similar backbones with complementary chromophore groups (ii) in which A and B have similar backbones with complementary functional groups and (iii) in which A and B have related functional groups but complementary backbones. Preferred LB films contain the compound (I).



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LANGMUIR BLODGETT FILMS

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10 This invention relates to Langmuir-Blodgett (LB) films, and to devices using such films.

An LB film is generally prepared by depositing a monolayer of a surface-active material onto a water surface, using now well established techniques. The molecules of the surface active material align in the monolayer, the hydrophilic ends remaining in the water, and the hydrophobic end projecting out of the surface. By other known techniques, this monolayer may be transferred essentially intact onto the surface of a solid substrate, and further monolayers deposited on the layer on the substrate to form a film, ie a
20 Langmuir-Blodgett film.

Compounds used in LB films are generally long chain molecules, containing 14 or more carbon atoms in their chain, with a hydrophilic "head" and a hydrophobic "tail", such as long chain carboxylic acids, eg $\text{CH}_3(\text{CH}_2)_{14}\text{CO}_2\text{H}$, where the CO_2H group constitutes
25 the hydrophilic head.

LB films have properties, which make them useful in certain electronic devices.

Conventional (Y-type) LB films are symmetrical about their layer planes; However one of the most interesting features of the
30 LB technique is the ability it gives to carry out the assembly of complicated molecular systems by depositing successive layers of different materials². In particular multilayer films of the general type ABAB.. where A and B are monolayers of different materials, and are non-centrosymmetric, are expected to show interesting and potentially useful pyroelectric, piezoelectric and
35 second order non-linear optical effects. A two compartment LB trough has been developed which allows the convenient fabrication of

ABAB.. films³.

Interest in pyroelectrics for thermal detectors is generally increasing because of their wide spectral response, sensitivity and lack of cooling requirements. Thin film pyroelectrics in particular are provoking much interest due to the shorter response times available with low thermal mass, large area capability, and compatibility with electronic readout devices such as the charge coupled device (CCD) or the pyro-FET. Being a low temperature technique the LB technique is compatible with semiconductor materials and large areas of uniform deposition are in principle possible.

To produce a pyroelectric effect a material must possess a spontaneous polarisation P_s ($C\ cm^{-2}$) thus it must also possess a polar axis. This requires that the material be non-centrosymmetric. Normally P_s is not directly observable due to compensation of the surface polarization charge by conduction charges from the material or atmospheric ions. However P_s may be a function of the temperature T , and if T is changed slightly the compensation charges are too slow (since pyroelectrics are dielectrics) to follow the change in P_s and a charge $\pm Q$ appears on the surfaces perpendicular to the polar axis; this is the pyroelectric effect.

We define the pyroelectric coefficient by

$$P = \frac{dP_s}{dT} \ (C\ cm^{-2}\ K^{-1})$$

and therefore

$$Q = p\Delta T$$

Generally p is also $p(T)$ but the variation is normally weak except near phase transitions in ferroelectric pyroelectrics. Therefore we can write $Q = p\Delta T$ where ΔT is a small change in T . Conventional high performance pyroelectrics have $p = 3 \times 10^{-8}\ C\ cm^{-2}\ K^{-1}$. Nearly all such materials are ferroelectrics operating close to their Curie temperature to obtain the high dP_s/dT .

In conventional Y-type LB films the molecular dipole moments of adjacent layers cancel and therefore $P_s = 0$. In an ABAB ... film the dipole moments of A and B being different do not cancel but add vectorially to give a non zero P_s and a polar axis perpendicular to the film plane. X and Z type films will also show a non-zero P_s ; pyroelectricity and piezoelectricity in such films has been reported

by Blinov^{4,5}.

It is an object of the present invention to provide novel LB films which have improved pyroelectric and piezoelectric properties. Other objects and advantages will be apparent from the following description.

According to a first aspect of the invention a Langmuir-Blodgett film ofABAB..... structure is composed of compounds A and B which are different, A and B containing complementary groups that result in non-centrosymmetry of the layer plane, the layers of A and B being optionally separated by a counter layer.

By "complementary groups" it is meant that A and B contain different atomic groupings in their structure so as to result in non-centrosymmetric layering in the LB film.

Compounds A and B will normally be long chain molecules containing 14 or more carbon atoms, in an essentially linear skeleton, one end of the molecule being of generally hydrophilic character, the other being generally hydrophobic.

In preferred embodiments of the invention, the compounds A and B may have the following general structural types (i), (ii) and (iii):

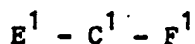
(i) Compounds A and B may have a generally similar backbone in terms of disposition of functional groups, the length of their backbone, their molecular shape and polarity, but wherein A and B contain groups which are chromophores, which may be the same or different, which result in non-centrosymmetric layering of A and B in the LB film.

In one embodiment of this type, the chromophore group in A and B is the same, but the group is disposed within their molecular structures rotated through 180° about an axis perpendicular to the molecular skeleton relative to the positions of the hydrophilic and hydrophobic ends.

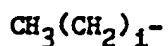
For example A may have a structure:



and B may have a structure:

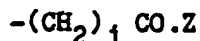


where E and E¹ may be the same or different, and are independently selected from:



where i is 2-20,

F and F^1 may be the same or different, and are independently selected from:



- 5 where j is 0 to 10, preferably 2 to 4 and Z is selected from $-OH$, C_{1-6} alkyl and NR^1R^2 where R^1 and R^2 may independently be H or C_{1-6} alkyl. F is a hydrophilic group. E, E^1, F and F^1 may contain the same or different numbers of carbon atoms, in each case the total in E plus F and E^1 plus F^1 being 14 or more.

- 10 C and C^1 are chromophores which result in non-centrosymmetric layering of A and B in the LB film. C and C^1 may be the same or different, and are preferably selected from



- 15 where X and Y are each independently selected from:

$=NH$, $=NR^3$ (R^3 being C_{1-4} alkyl), $=CO$, $=CH(CN)$, $=CH(ONO_2)$, $=CH$ (Halogen), $-COO-$, $-OOC-$, $=CH(ONO)$, $-CO-NH-$, $-CO-NR^4-$ (where R^4 is C_{1-6} alkyl) and $-O-$.

- 20 It is preferred to use polar groups X and Y which undergo solid state reorganisation, to give a large dP_s/dT in the region of room temperature.

C and C^1 may respectively be $X-\text{C}_6\text{H}_4-Y$ and $Y-\text{C}_6\text{H}_4-X$ in the molecules A and B , ie to give structures of the type:



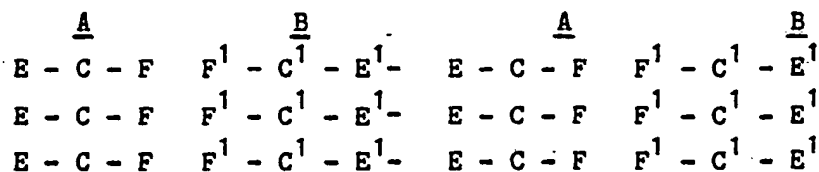
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and

so that in these structures the chromophore is rotated through 180° relative to the positions of the hydrophilic and hydrophobic ends.

- 30 In LB films of the type (i) non-centrosymmetric layering occurs as below:



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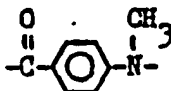
(the layer planes here being shown as running vertically)

In LB films of type (i), if it is considered necessary to

interpose one or more counter-layers between the layers AB, a preferred material for counter-layers is a long chain alkyl amine of formula $\text{CH}_3(\text{CH}_2)_n\text{NH}_2$ where n is 16 to 26, preferably 18 to 20.

A particularly preferred chromophore group is

5



(ii) Compounds A and B may have a generally similar backbone in terms of the length of their backbone, their molecular shape and polarity, but in which A and B contain different complementary functional groups which result in non-centrosymmetric layering in the LB film.

The functional groups which result in non-centrosymmetric layering are preferably located at the terminal position of the molecule.

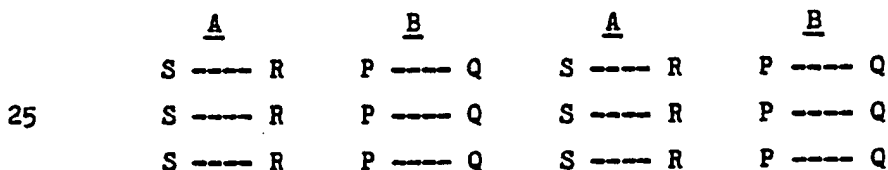
For example A may be a molecule:



and B may be a molecule:



where R and P may be the same or different functional groups or hydrophobic termini, and Q and S are different functional groups, ---- representing the molecular backbone, non-centrosymmetric layering in the LB film occurs as below:



(the layer planes being shown as running vertically)

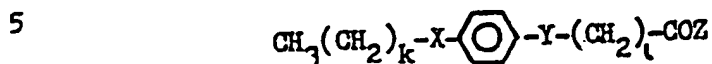
Various preferred embodiments of this type are discussed below:

30 (a) A is a long chain alkyl amine of the structure $\text{C}_n \text{H}_{2n+1} \cdot \text{NH}_2$ where n is an integer between 16 and 26, preferably 18 to 20. It has been found that at n greater than 24 the chain begins to become too rigid for use at room temperatures, and that at n less than 16 the amine becomes too soluble in the subphase.

35 B is a long chain fatty acid of the structure $\text{R}^5 (\text{CH}_2)_m \text{COOH}$ where R^5 is selected from CH_3 and $\text{CH}_2=\text{CH}$, and where m is indepen-

dantly an integer within the limits and preferences set above for n.

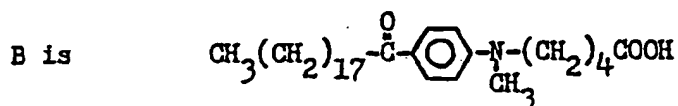
(b) A is a long chain alkyl amine as defined in (ii)(a) above, and B is a compound of general structure:



in which k is 2 to 20, l is 0 to 10 preferably 4, X, Y and Z are as defined in connection with type (i) above.

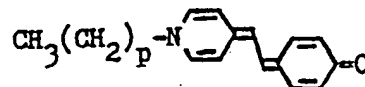
An example of such an A-B combination is

10 A is stearylamine (n-octadecylamine $\text{C}_{18}\text{H}_{37}-\text{NH}_2$)



(c) A is a long chain alkyl amine as defined in (ii)(a) above, and

15 B is a compound of the structure:



where p is 6 to 25, preferably no more than 21.

(iii) Compounds A and B may have the same or a closely related functional group and different complementary structural backbones which result in non-centrosymmetric layering in the -ABAB- film.

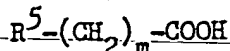
The functional group is preferably located at a terminal position in the molecule. A preferred functional group is -COZ, with Z as defined above, especially -OH ie a carboxylic acid group. One or more of the compounds A and B may contain a chromophore, such as



where X and Y are as defined above.

In a preferred LB film of this type:

A is a long-chain alkyl carboxylic acid of general formula:



where R^5 and m are as defined above, and

B is a chromophore-containing compound of general formula:



where k, X, Y, l and Z are as defined above.

In this description the term "chromophore" is used as that part of a molecule which produces its colour, or its light absorbance in the non visible region of the spectrum. Preferred chromophores are polar chromophores, as it is polarity of the chromophore which is responsible for their pyroelectric activity in LB films rather than their light absorbance.

Langmuir Blodgett films according to the invention show pyroelectric activity in the region of room temperature. Some may also show piezoelectric behaviour. Films according to the invention which have a high degree of polarisable structure may also show non-linear optical properties.

Films according to the invention also have the advantage that their polarisation is solidly built into their structure, and is thus unlikely to be lost or reversed at an electric field strength below the dielectric breakdown field.

Langmuir Blodgett films according to the invention may be prepared by known means, for example using the apparatus and method described in reference 3.

Films according to the invention may be used in electro optical and other electronic devices which use a Langmuir Blodgett film as a sensing element, and therefore according to a further aspect of the invention an electronic device is provided which incorporates a Langmuir Blodgett film as described above as its sensing element. In particular there is provided a thermal sensor which incorporates as its sensing element a Langmuir Blodgett film according to the invention. Such devices may be constructed using methods known in the art.

The invention will now be described by way of example only with reference to the accompanying figures 1 to 4, which show:

Fig.1 Voltage responsivity versus chopping frequency for ...ABAB... LB film of stearylamine / ω - tricosenic acid. Total thickness 500 layers.

Figs. 2, 3, 4 Pyroelectric coefficient (p) versus absolute temperature ($^{\circ}\text{K}$) for various LB films of types (i), (ii)(a), (ii)(b) and (iii).

Experimental Details

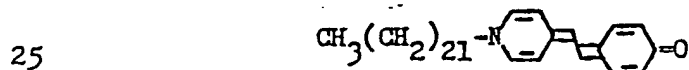
1. Materials

Saturated carboxylic acids were obtained from Sigma (99%).

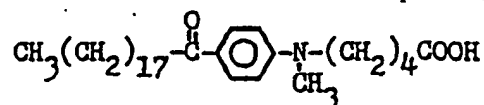
15 Stearylamine (n-octadecylamine) was obtained from ICN Pharmaceuticals Inc. ω -tricosenoic acid ($\text{CH}_2=\text{CH}(\text{CH}_2)_{20}\text{COOH}$) was kindly supplied by Dr A Barraud and A Ruaudel-Teixier (CEN Saclay) and Gedye₂₂ was kindly supplied by Dr G L Gaines Jr., (General Electric Company). All the above were used without further purification.

20 The dye called 18ON4A was synthesised using a standard method. All the materials were spread from high purity solvents (BDH Ltd Aristar grade) usually 1,1,1, trichloroethane or chloroform.

Gedye₂₂ is an abbreviation for the compound:

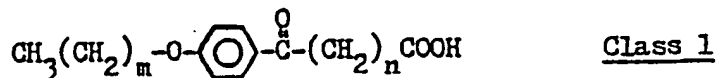


18ON4A is an abbreviation for the compound:



Other compounds were synthesised as described below.

5 Compounds of the general formula:

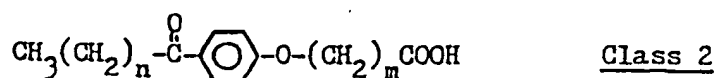


(a) Phenol (0.1 mole) and sodium hydroxide (0.1 mole) in ethanol
10 (100 ml) were stirred and refluxed during the dropwise addition of
an appropriate 1-bromo alkane (0.095 mole), stirring and refluxing
being continued until the 1-bromo alkane was consumed, as evidenced
by TLC (typically 12 hr) after removal of the ethanol solvent, the
reaction mixture was diluted with water and the product extracted
15 with chloroform. The dried extract was vacuum evaporated and the
residue crystallised from light petroleum to give the phenoxy alkane
(Yield ca. 80%).

(b) Aluminium chloride (0.03 mole) was added to a solution of the
above phenoxy alkane (0.01 mole) and the methyl ester chloride of
20 the appropriate alkane dioic acid (0.0105 mole) in dichloromethane
(20 ml). The mixture was stirred at room temperature for 24 hr or
until TLC showed the absence of starting material, then hydrolysed
with dilute hydrochloric acid and ice. Chloroform extraction and
vacuum evaporation gave the keto-ester which was recrystallised from
25 light petroleum (Yield ca. 80%).

(c) The above keto-ester (0.01 mole) was refluxed with sodium
hydroxide (0.04 mole) in ethanol (15 ml.) and water (2 ml.), until
TLC showed the absence of starting material (ca. 1 hr). After removal
of the solvent ethanol the product was collected by chloroform
30 extraction. Concentration of the dried extract gave the ketoacid
which was further purified by recrystallisation from chloroform.

Compounds of the general formula:



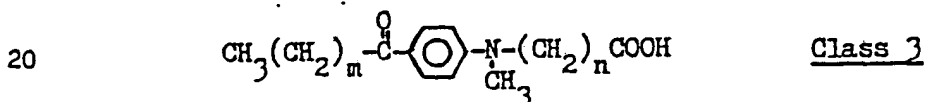
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(a) w-Phenoxyalkanoic esters. The requisite w-bromalkanoic acid
ester was made by stirring the commercially available acid (2g.)

with the appropriate alcohol (70 ml.) at room temperature, the reaction being catalysed by the addition of a small amount of acetyl chloride. After 24 hr., or when TLC indicated total conversion, the excess alcohol was removed by vacuum distillation and the ester taken up in chloroform. The extract was washed with sodium bicarbonate, water, then dried and vacuum evaporated to give the ester which was used directly without further purification. Alkylation of phenol with the w-bromoester was carried out as in 1(a) above using the same alcohol as solvent as was used in the previous esterification step to avoid trans-esterification.

(b) Acylation of w-phenoxyalkanoic esters. The requisite acyl chlorides were prepared from commercially available alkanolic acids by reaction with thionyl chloride, and purified by vacuum distillation. The acylation reaction was carried out using the method outlined in 1(b) above, the resulting ketoester being hydrolysed to the w-(4-alkanoylphenoxy) alkanolic acid by the method of 1(c) above.

Compounds of the general formula:



(a) w-(N-methyl-N-phenylamino) alkanolic esters. A mixture of N-methylaniline (0.1 mole), the methyl or ethyl w-bromoalkanoate of 2(a) above (0.1 mole), and sodium carbonate (0.3 mole) in hexamethylphosphorus triamide (50 ml.) was stirred and heated at 140°C until TLC showed total conversion of starting material (ca. 20 hr.), then cooled and diluted with water (200 ml.). The precipitated acid was collected, washed, and recrystallised from ethanol. Acids which precipitated as oils were collected by chloroform extraction prior to recrystallisation.

(b) w-N-(4-formylphenyl)-N-methylamino alkanolic esters. The aminoester from (a) (0.005 mole) was added to phosphorus oxychloride (5 ml.) in N,N-dimethylformamide (20 ml.) at room temperature. The mixture was heated at 90-95°C for 2 hr., then cooled and neutralised with aqueous sodium hydroxide. The crude aldehydic ester was collected by extraction with ethyl acetate and purified by chromatography over silica in ethyl acetate - light petroleum, prior to

recrystallisation from light petroleum (Yield ca. 75%).

(c) A solution of Grignard Reagent was prepared from the requisite alkyl halide (0.006 mole) and magnesium turnings in dry tetrahydrofuran (30 ml.) under argon. The above aldehydic ester (0.004 mole) in dry tetrahydrofuran (5 ml.) was cooled to -15°C and the solution of Grignard reagent added dropwise during 10 min. The reaction mixture was stirred at $-15 - -5^{\circ}\text{C}$ until TLC showed no further conversion (ca. 45 min), then hydrolysed with aqueous ammonium chloride and extracted with ether. Evaporation of the dried ethereal extracts and chromatography of the residue over silica in ethyl acetate - light petroleum gave the alcohol (ca. 55%) which was oxidised directly.

(d) The above alcohol in dichloromethane (5 ml.) was treated with pyridinium dichromate (0.014 mole) and stirred at 0°C for 12 hr. The reaction mixture was diluted with water and the ketoester collected by ether extraction. It was purified by chromatography over silica with ethyl acetate - light petroleum (Yield ca. 30%).

(e) The ketoester (0.003 mole) was refluxed with sodium hydroxide (0.009 mole) in ethanol (18 ml.) and water (2 ml.) until reaction was complete as determined by TLC. On cooling, the solution was acidified with dilute hydrochloric acid. The precipitated ketoacid was collected, washed and recrystallised from chloroform (Yield ca. 70%).

Using the methods described above, compounds as follows were prepared:

	<u>m</u>	<u>n</u>
<u>Class 1</u>	8	4
"	11	4
"	15	4
30 "	15	2
<u>Class 2</u>	17	4
<u>Class 3</u>	17	4
"	17	6
"	17	8

35 2. Multilayer Deposition

All films were deposited on aluminized glass cover slips cut to 2 cm square. The cover slips were cleaned ultrasonically in Decon-

90 followed by rinsing in pure filtered water and drying in a stream of nitrogen before evaporation of the 1 cm wide aluminium electrode stripe ($> 60\text{nm}$ thickness) in a cryogenically pumped (oil free) vacuum coater. Thickness control was achieved by use of a quartz crystal monitor at the substrate plane. The source was a resistively heated tungsten filament and the source to substrate distance was 40 cm.

Stearylamine/carboxylic acid ABAB... films were produced using the RSRE two-compartment trough³. Each compartment was maintained at a surface pressure of 30 mN m^{-1} over a subphase of pure water (original resistivity $18 \text{ M}\Omega \text{ cm}$). Transfer to the substrates was effected at $\sim 5 \text{ mm sec}^{-1}$ with the amine being deposited on the down stroke. Transfer ratios were near unity.

Stearylamine/Gedye₂₂ films were deposited with the surface pressure in the Gedye compartment being 25 mN m^{-1} . The subphase contained 10^{-3} M ammonium sulphate and was maintained at a pH of 7.5 ± 0.3 . Transfer was effected at $\sim 1 \text{ cm min}^{-1}$ to allow the Gedye film to drain. Stearylamine was transferred on the down stroke.

Stearylamine/18ON4A films were deposited under identical conditions to the stearylamine/carboxylic acid films.

In all cases two substrates were coated together. After deposition the films were left for 24 hrs before evaporation of a top aluminium electrode stripe perpendicular to the first. The final device area was 1 cm^2 .

3. Preliminary Film Assessment

(a) Infrared spectroscopy

Infrared spectra were obtained using a Perkin Elmer 983G infrared spectrometer. The samples were deposited as films 50 layers thick on the polished faces of Ge prisms ($5 \times 1.5 \times 0.2 \text{ cm}$, Harrick Scientific Corp) and the measurement was made using the technique of attenuated total reflection (ATR)⁶.

(b) Dielectric Measurement

Capacitance and $\tan \delta$ data were obtained using a Hewlett-Packard 4274ALCR meter. The signal amplitude was set at 50 mV. All measurements were made in laboratory ambient conditions on the samples assessed for pyroelectric activity.

(c) X-ray d-spacings

Measurements were made on a 50 layer ABAB film of stearyl-amine / arichidic acid deposited on a silanized unaluminized cover slip. A low - angle diffractometer with Cu K radiation was used.

5 4. Pyroelectric Assessment.

This was carried out using dynamic and static methods as described in G W Smith, M F Daniel, J W Barton and N Ratcliffe, Proc. Conf. on Langmuir Blodgett Films, Schenectady NY, July 1 - 4 1985, in Thin Solid Films (TSF ref. 8560).

10 Results

1. Infrared Spectra.

The results indicated the formation of the ions $C_{19}H_{39}COO^-$ and $C_{19}H_{37}NH_3^+$ due to proton transfer between the acid and the amine. The 'salt' of these two ions has therefore been formed,
15 with a large dipole moment.

2. Pyroelectric Measurements.

The static method was found to allow determination of temperature changes more accurately and results are summarised briefly in Table 1.

20 In Table 1 the following values are given:

d = X-Ray d- spacings

ϵ_r = dielectric constant

$\tan \delta$

p = pyroelectric coefficient

25

Since P_s is dipole moment/unit volume and since the hydrocarbon tails of the saturated fatty acids contribute to the molecular volume but not significantly to the dipole moment we attempted to introduce a systematic variation in P_s and hence hopefully in p by varying the length of the fatty acid B layers in our films. It is obvious from Table 1 that this objective has not been achieved. Calculation yields an expected change in P_s of 15% on going from C_{18} acid to C_{24} acid and this small change, if it appears at all in p , is lost in the scatter of the results. All the amine/saturated acid films give $p \sim 5 \times 10^{-11} \text{ C.cm}^{-2}.\text{K}^{-1}$.

Films of stearylamine with the unsaturated fatty acid ω -tricosenic acid consistently gave values twice as high $p = 1 \times 10^{-10} \text{ C.cm}^{-2}.\text{K}^{-1}$. This presumably reflects the temperature dependence of the dipole moment of the terminal double bond (0.3 D) since that is the only difference between ω -tricosenoic acid and saturated carboxylic acids.

Films incorporating the merocyanine dye Gedye₂₂ gave very low values of p ($8 \times 10^{-12} \text{ C cm}^{-2} \text{ K}^{-1}$). These films when measured had the dye in its yellow form² with the enolic oxygen protonated, which reduces the intramolecular charge transfer contribution to the dipole moment. Also there is no amine/acid salt formation. So P_s for these films may be small compared with the amine/acid films and this may explain the small value of p . The higher value for ϵ_r is an indication of the high polarisability of the chromophore which has possibilities for non-linear optics.

Also included in Table 1 for comparison are the data obtained for a Y-type film of ω -tricosenoic acid showing the very small observed pyroelectric effect in a centrosymmetric LB film which may be due to small differences in the transfer ratios in the up and down directions or residual drift in the apparatus.

The temperature dependence of p for ABAB... films of stearylamine with various fatty acids is shown in Figure 2. The low

Table 1.

	A	B	d (μm)	ϵ_r (1kHz)	$\tan \delta$	$p(\text{Ccm}^{-2}\text{K}^{-1})$ (280-285K)
5						
	$\text{CH}_2\text{CH}(\text{CH}_2)_{20}\text{COOH}$	$\text{CH}_2\text{CH}(\text{CH}_2)_{20}\text{COOH}$	0.45	2.85	0.0007	$2 \times 10^{-12} *$
	$\text{C}_{19}\text{H}_{37}\text{NH}_2$	$\text{C}_{17}\text{H}_{35}\text{COOH}$	0.26	2.71	0.016	3.2×10^{-11}
10	..	$\text{C}_{19}\text{H}_{39}\text{COOH}$	0.13	2.48	0.020	5.8×10^{-11}
	..	$\text{C}_{21}\text{H}_{43}\text{COOH}$	0.23	2.82	0.033	3.3×10^{-11}
	..	$\text{C}_{23}\text{H}_{47}\text{COOH}$	0.15	2.66	0.009	6.0×10^{-11}
	..	Gedye ₂₂	0.23	6.49	0.04	8.1×10^{-12}
15	..	18ON4A	0.14	3.24	0.006	2.25×10^{-10}
	..	$\text{CH}_2\text{CH}(\text{CH}_2)_{20}\text{COOH}$	0.40	2.91	0.094	1.14×10^{-10}

Gedye₂₂ is: CC(C)(C)CCCCCCCCCCCCCCCCCCCCC1=CC=CC=C1C=C2C=CC(=O)C=C2

18ON/4A is: $\text{CH}_3(\text{CH}_2)_{17}-\overset{\text{O}}{\underset{\text{||}}{\text{C}}}-\text{C}_6\text{H}_4-\underset{\text{CH}_3}{\underset{|}{\text{N}}}-(\text{CH}_2)_4\text{COOH}$

* Y-type film for comparison.

temperature limit was determined by the cooler and the high temperature limit by the background current, which varied from film to film but always increased sharply above room temperature. it is obvious that p increases at higher temperatures with the rate of increase being much higher above about 280K. Below that temperature all three curves are quite parallel. This general trend of $p(T)$ was found for all films examined.

X-RAY MEASUREMENTS

The bilayer repeat distance for a 50 layer film of stearylamine arachidic acid has been determined to be 49 \AA . This is somewhat smaller than the value one would calculate from measuring molecular models (54 \AA) and indicates a degree of molecular tilt within the layer planes.

DISCUSSION

By making certain assumptions it is possible to calculate an approximate maximum value of P_s for the amine/acid films. Assuming complete proton transfer, a resulting charge separation of 0.3 nm and a bi-molecular (acid + amine) volume of $1.25 \times 10^{-27} \text{ m}^3$ we obtain $P_s = 4 \text{ \mu C cm}^{-2}$. The observed values for p therefore represent a change in P_s of $\sim 10 \text{ ppm K}^{-1}$. It is interesting to note that the calculated value for P_s is similar to that measured for the high performance pyroelectric TGS (triglycine sulphate)⁹ ($2.7 \text{ \mu C cm}^{-2}$). The greater than 500 fold difference in p is therefore even more striking. However TGS is ferroelectric which means that its polarisation can be switched by a sufficiently high field applied for long enough and that its polarisation vanishes above its Curie temperature by a crystallographic rearrangement. For an ABAB... film the polarisation is solidly built into the structure and one cannot conceive of any electric field short of the dielectric breakdown field which could change that. Similarly no temperature below the melting point is likely to change P_s significantly. The low values of p are thus explained by the fact that the charge on the headgroups is probably invariant with temperature and the only change in P_s is due to a change in molecular volume due to thermal expansion.

Films containing the 18 ON 4A dye will have larger values of P_s than stearylamine/fatty acid films because the dipole moment of the

chromophore is in the same sense to that at the head group. The higher observed value of p then probably means one of two things. Either the dipole moment or the orientation of the chromophore within the films changes thus changing the projection of its dipole moment on the polar axis (film normal) and its contribution to P_s . We cannot as yet decide between these two options.

In all the films so far examined the effect of increasing the temperature is to decrease P_s .

CONCLUSIONS

10 We have succeeded in producing ABAB... films of stearylamine with various film forming materials. As expected these films exhibit pyroelectric activity and we have measured this as a function of temperature and material. Calculation of P_s for the amine/fatty acid films yields a surprisingly high value which is not reflected in p . Films containing 18 ON 4A, display a larger value for p . It thus appears that much can be done to increase p by introducing polar groups with the right characteristics.

In a further series of experiments, LB films of types (iii) and (i) above were prepared and investigated. These contained the compounds:

180N4A (see above)

5

18NO4A

$$\text{CH}_3(\text{CH}_2)_{17}\text{N}-\text{C}_6\text{H}_4-\overset{\text{O}}{\overset{\text{||}}{\text{C}}}-(\text{CH}_2)_4\text{COOH}$$

140' 04A

$\text{CH}_3(\text{CH}_2)_{13}\text{O}-\text{C}_6\text{H}_4-\overset{\text{O}}{\underset{\text{||}}{\text{C}}}\cdot(\text{CH}_2)_4\text{COOH}$

10

Eicosanoic acid $\text{CH}_3(\text{CH}_2)_{18}\text{COOH}$

Experimental Method.

The compounds were repeatedly recrystallised and their purity checked by thin layer chromatography and mass spectroscopy.

Dilute solutions (0.2 - 1.0 mg/ml) in 'Aristar' (BDH Ltd, UK) grade solvents were spread on a commercial two compartment trough incorporating the RSRE gate design as described above. Pressure was maintained at $\pm 1 \text{ mN m}^{-1}$ of the set transfer pressure for each material. The subphase conditions were $\text{pH} = 8.5 \pm 0.2$; $\text{Cd}^{2+} = 2.5 \times 10^{-4} \text{ M}$. The subphase had an original resistivity of $18 \text{ M } \Omega \text{ cm}$. Substrate speeds of 1 mm/sec were used. 18ON4A was normally dipped on the up stroke, 14O' 04A and 13NO4A on the up stroke. Eicosanoic acid could be dipped on either stroke. The substrates and electrode arrangement was as described above.

Pyroelectric coefficients and dipole moments were determined by known techniques (ref. 10 and 11). An ambient of dry nitrogen gas was used throughout.

Results.

~~(1) Type (iii) Films.~~

Results are given in Table 2 below, and graphically in Fig 4. The sign of the pyroelectric current was the same in all three films. Cooling produced a current from the bottom to the top electrode, therefore Ps increases with decreasing temperature. Reversing the deposition order for the 18ON4A / eicosanoic acid film reversed the direction of current flow. p increased with temperature for most of the films.

Table 2

5	A	B	μ ($\times 10^{-9} \text{Cm}$)	bilayer rep. dist.	tilt $^{\circ}$	Ps ($\times 10^{-3} \text{Cm}^{-2}$)	p ($\times 10^{-7}$)
	eicosanoic acid	140' O ₄ A	1.2 \pm 0.3	5.52nm	28.9	7.9	1.6
	..	18ON ₄ A	1.9	5.4nm	43	9.3	1.9

10 note: tilt = tilt of chromophore containing group B with
respect to film normal.

In these films use of 18ON₄A is thus shown to give better
Ps and p.

15 (2) Type (i) films.

Two different structures of type (i) containing a chromo-
phore in each layer have been prepared, ie:

	<u>A</u>	<u>B</u>
	140' O ₄ A	18ON ₄ A
and	18NO ₄ A	18ON ₄ A

20 The results for these two films are shown in Fig. 3.

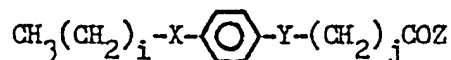
The pyroelectric response of 140' O₄A / 18ON₄A was about
twice that of 18NO₄A / 18ON₄A.

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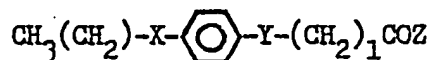
CLAIMS

1. A Langmuir-Blodgett (LB) film of ABAB structure composed of compounds A and B which are different and which contain complementary groups that result in non-centrosymmetry of the layer plane the layers A and B optionally being separated by a counter-layer.
2. An LB film according to claim 1, characterised in that compounds A and B have a generally similar backbone in terms of disposition of functional groups, length, molecular shape and polarity and in which A and B contain chromophores, which may be the same or different, which result in non-centrosymmetric layering of A and B in the layer plane.
3. An LB film according to claim 2, characterised in that the chromophore group in A and B is the same, but the group is disposed within the molecular structures of A and B rotated through 180° about an axis perpendicular to the molecular skeleton relative to the positions of the hydrophilic and hydrophobic ends.
4. An LB film according to claim 3, characterised in that A and B are independently selected from compounds having a structure:



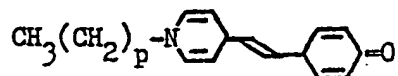
where i is 2-20, j is 1-10, z is selected from -OH, C₁₋₆ alkyl and NR¹R² where R¹ and R² may independently be H or C₁₋₆ alkyl, and where X and Y are independently selected from =NH, =NR³ (R³ being C₁₋₄ alkyl), -CO-, -CH₂CN, -CHONO₂, -CH(Halogen), -COO-, -OOC-, -CH(ONO)₂, -CO-NH-, -CO.NR⁴ (where R⁴ is C₁ - C₆ alkyl) and -O-.

5. An LB film according to claim 1, characterised in that A is a long chain alkyl amine of structure C_nH(2n+1)-NH₂ where n is 16-26, and B is a long chain fatty acid of structure R⁵-(CH₂)_m-COOH where R⁵ is selected from CH₃ and CH₂=CH and where m is 16-26.
6. An LB film according to claim, characterised in that A is a long chain alkyl amine of structure C_nH(2n+1)-NH₂ where n is 16-26, and B is a compound of general structure:



in which k is 2-20, l is 0 to 10, and Y and Z are as defined in claim 4.7

7. An LB film according to claim 1 characterised in that A is a long chain alkyl amine of structure $\text{C}_n\text{H}(2n+1)\text{NH}_2$ where n is 16-26, and B is a compound of structure:



where p is 6-25.

8. An LB film according to claim 1, characterised in that A is a long chain carboxylic acid of general formula $\text{R}^5(\text{CH}_2)_m\text{COOH}$ where R^5 and m are as defined in claim 5 and B is a compound of general formula:



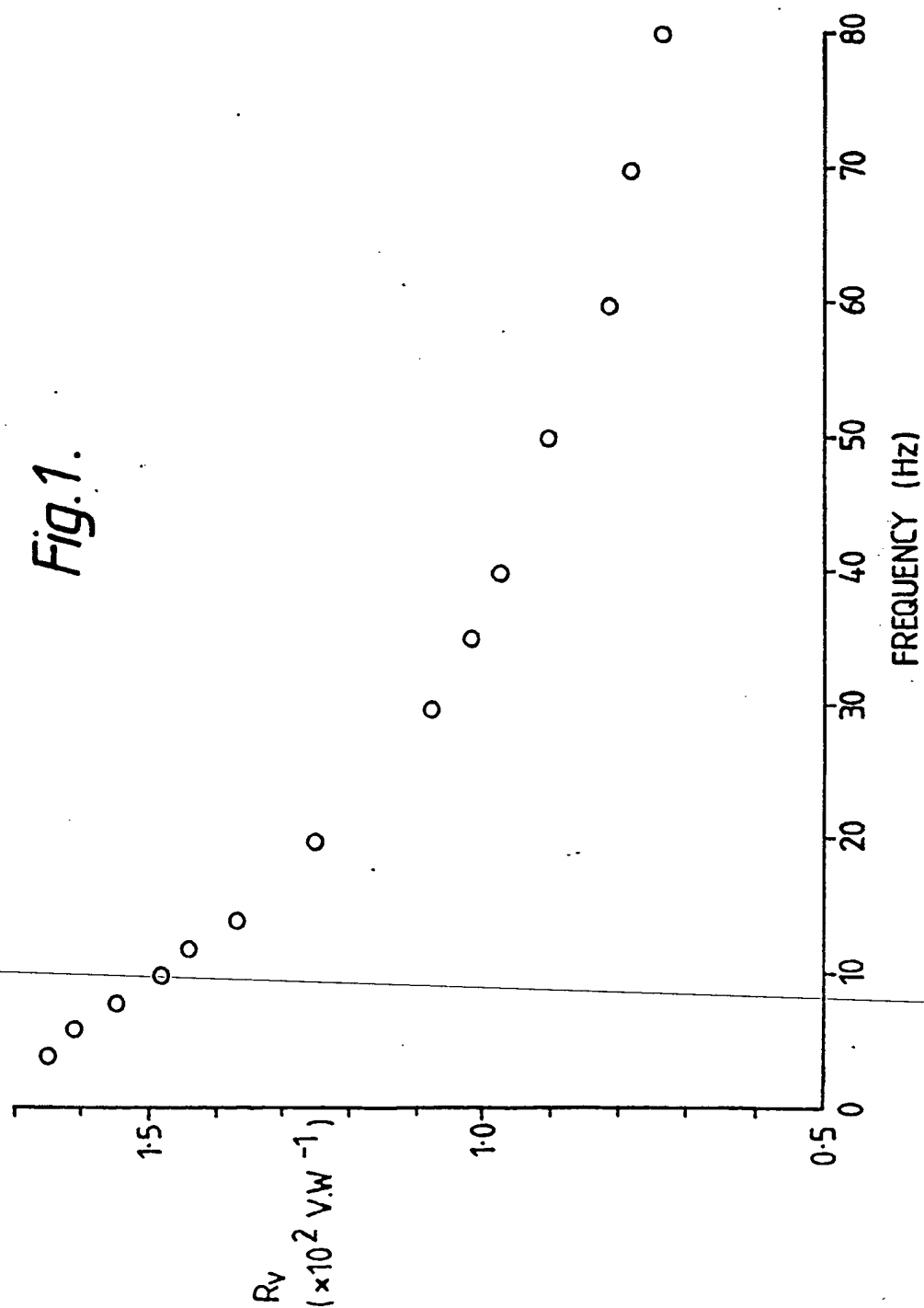
where k is 2-20, l is 0 to 10 and X , Y and Z are as defined in claim 4.

9. An LB film according to any one of claims 4, 6, or 8, in which X is $=\text{CO}$ and Y is $=\text{N}(\text{CH}_3)$ or X is $=\text{N}(\text{CH}_3)$ and Y is $=\text{CO}$.

10. An electro optical device incorporating an LB film according to any one of the preceding claims.

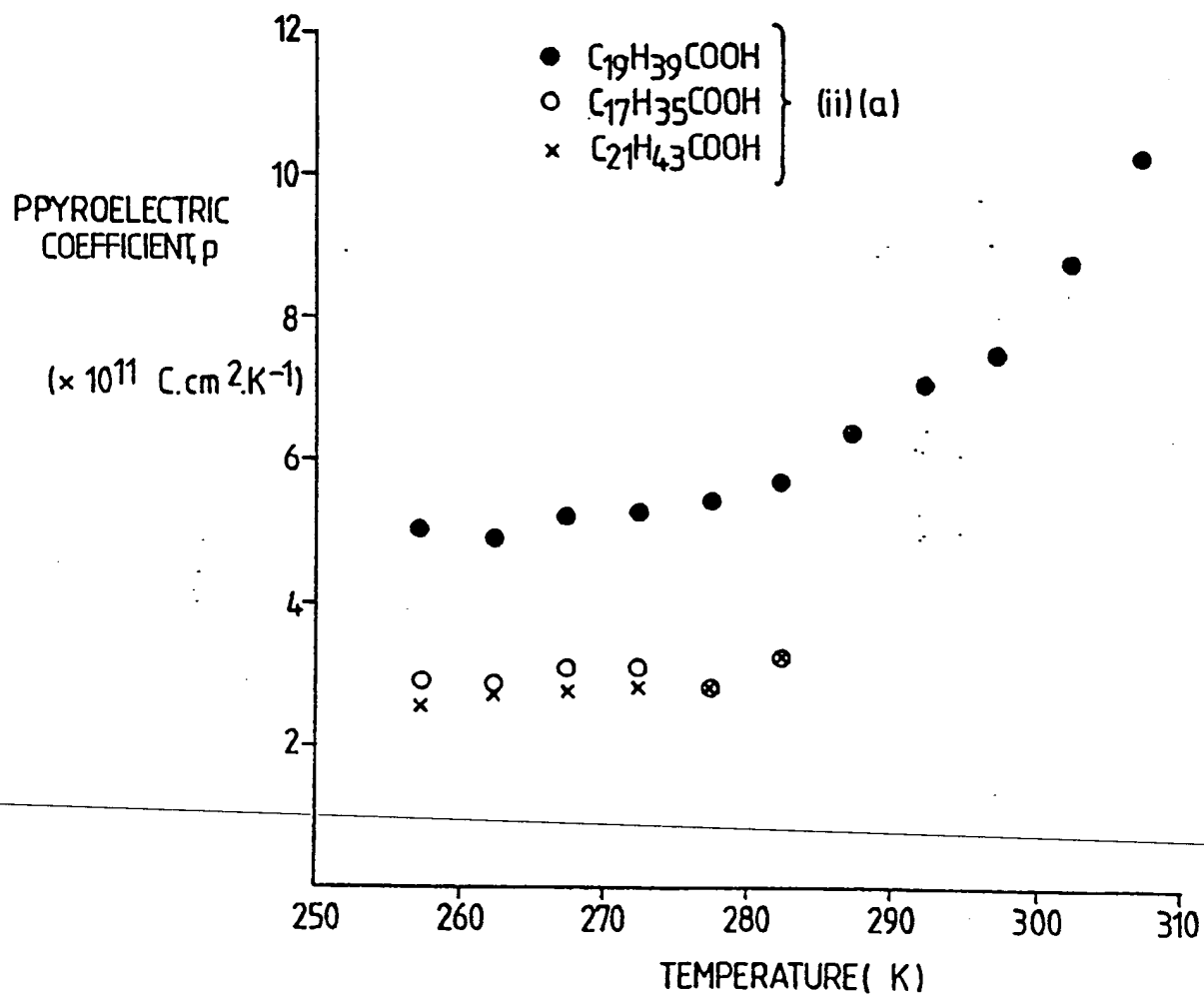
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Fig. 1.



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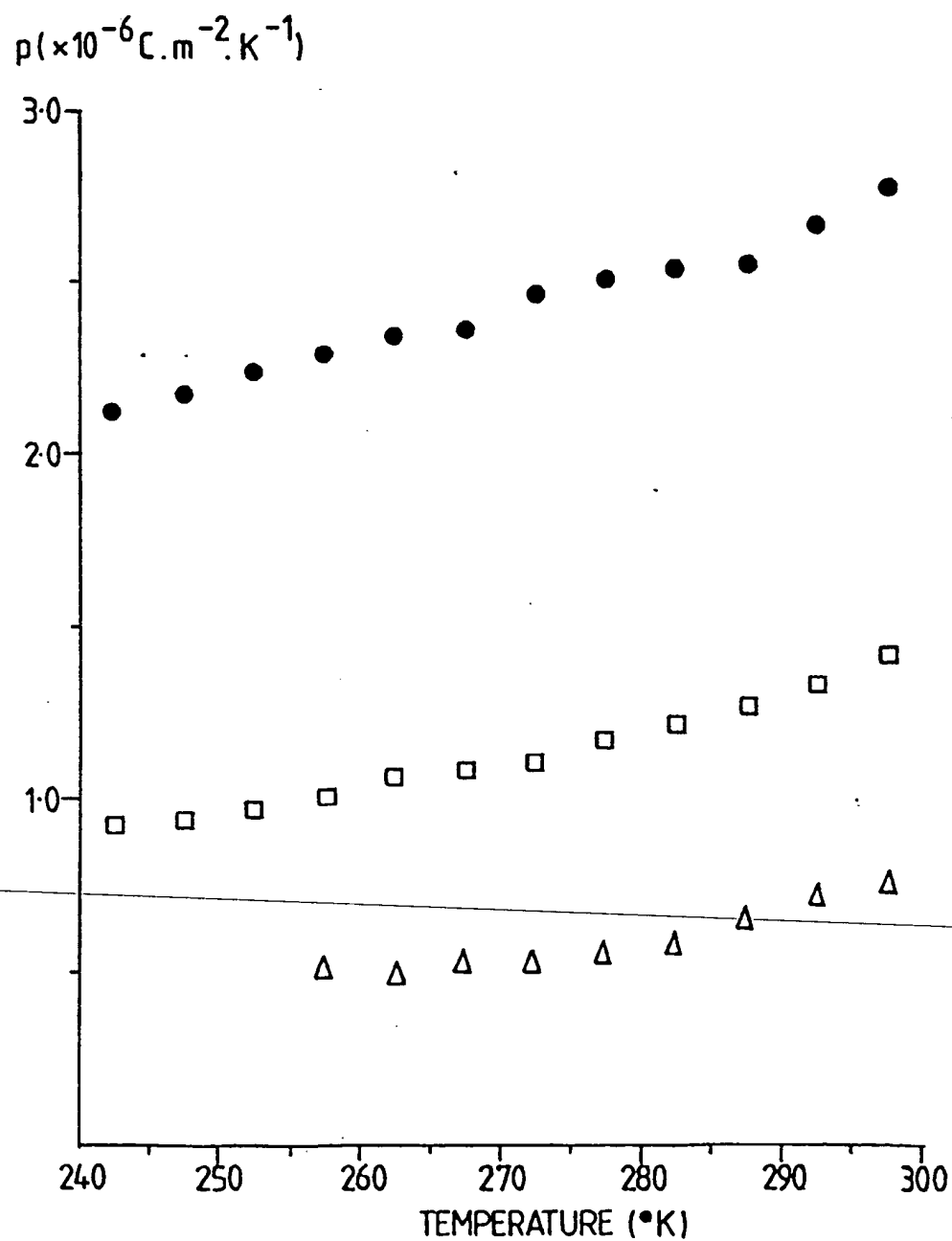
Fig. 2.



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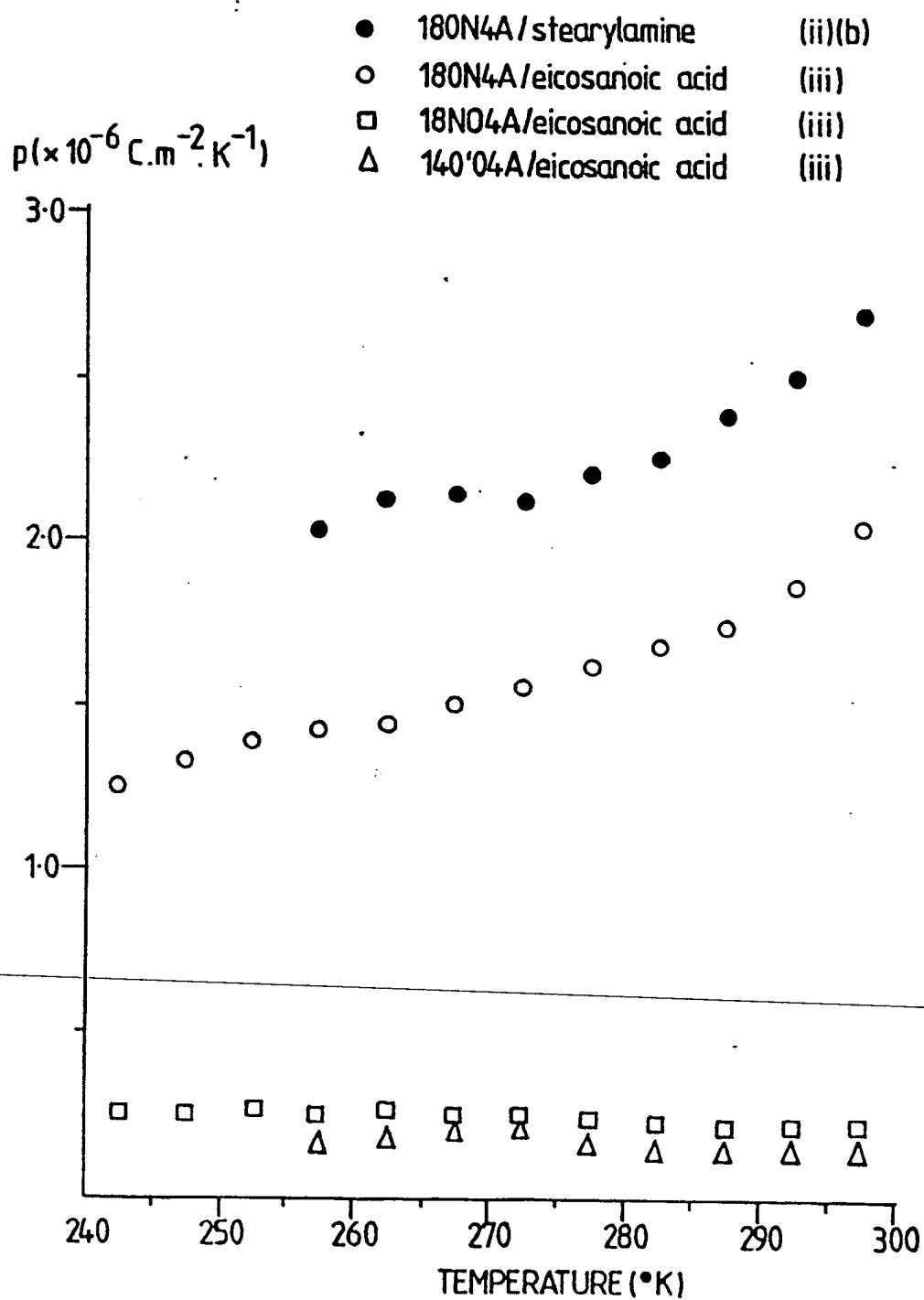
Fig. 3.

- 180N4A/140'04A (i)
- 180N4A/18N04A (i)
- △ eicosanoic acid / stearylamine. (ii)(a)



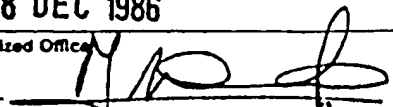
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Fig.4.



INTERNATIONAL SEARCH REPORT

International Application No PCT/GB 86/00371

I. CLASSIFICATION OF SUBJECT MATTER (If several classification symbols apply, indicate all) *		
According to International Patent Classification (IPC) or to both National Classification and IPC		
IPC ⁴ : H 01 L 29/28; H 01 L 37/02; H 01 L 41/18; G 02 F 1/37		
II. FIELDS SEARCHED		
Minimum Documentation Searched *		
Classification System	Classification Symbols	
IPC ⁴	H 01 L; G 02 F	
Documentation Searched other than Minimum Documentation to the extent that such Documents are included in the Fields Searched *		
III. DOCUMENTS CONSIDERED TO BE RELEVANT *		
Category *	Citation of Document, ** with indication, where appropriate, of the relevant passages ¹²	Relevant to Claim No. ¹³
X	Molecular Crystals and Liquid Crystals and Letters, volume 102, no. 617, 1984 (New York, US), M.F. Daniel et al.: "Preparation of non-centrosymmetric langmuir-blodgett films with alternating merocyanine and stearylamine layers", pages 193-198 see page 194, last paragraph - page 195, paragraph 2; page 197, paragraph 2 - page 198, paragraph 2 (cited in the application)	1,7
X	Thin Solid Films, volume 115, no. 1, May 1984, (Lausanne, CH), I.R. Girling et al.: "A method for the preparation of an alternating multilayer film", pages 85-88 see page 85, paragraph 2; page 86, paragraph 3 - page 87, paragraph 1; figure 5	1,2
A	Thin Solid Films, volume 120, 1984, (Lausanne, CH), L.M. Blinov et al.: "Polar langmuir-blodgett films", pages 161-170 see page 161, paragraph 1 - page 163, paragraph 1 (cited in the application)	1,4
<div style="display: flex; justify-content: space-between;"> <div style="width: 45%;"> <p>* Special categories of cited documents: ¹⁰</p> <p>"A" document defining the general state of the art which is not considered to be of particular relevance</p> <p>"E" earlier document but published on or after the international filing date</p> <p>"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)</p> <p>"O" document referring to an oral disclosure, use, exhibition or other means</p> <p>"P" document published prior to the international filing date but later than the priority date claimed</p> </div> <div style="width: 45%;"> <p>"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention</p> <p>"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step</p> <p>"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art</p> <p>"A" document member of the same patent family</p> </div> </div>		
IV. CERTIFICATION		
Date of the Actual Completion of the International Search		Date of Mailing of this International Search Report
10 November 1986		18 DEC 1986
International Searching Authority		Signature of Authorized Officer
EUROPEAN PATENT OFFICE		M. VAN MOL 

Category *	Citation of Document, with indication, where appropriate, of the relevant passages	Relevant to Claim No
A	EP, A, 0081477 (UNIVERSITY PATENTS INC.) 15 June 1983 see page 12, lines 1-21; page 40, lines 1-19; claims 1-3; figure 1 <div style="text-align: center;">-----</div>	

ANNEX TO THE INTERNATIONAL SEARCH REPORT ON

INTERNATIONAL APPLICATION NO. PCT/GB 86/00371 (SA 13760)

This Annex lists the patent family members relating to the patent documents cited in the above-mentioned international search report. The members are as contained in the European Patent Office EDP file on 25/11/86

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Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP-A- 0081477	15/06/83	EP-A,B 0021695	07/01/81
		EP-A,B 0022618	21/01/81
		JP-A- 56042229	20/04/81
		JP-A- 56043220	21/04/81
		AU-A- 5895880	08/01/81
		AU-A- 5903580	08/01/81
		EP-A,B 0077577	27/04/83
		CA-A- 1158598	13/12/83
		US-A- 4431263	14/02/84
		AU-B- 541737	17/01/85
		AU-B- 546234	22/08/85
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		US-A- 4536450	20/08/85